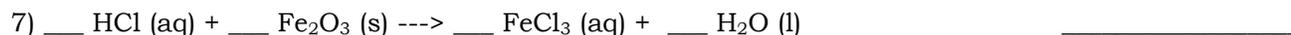
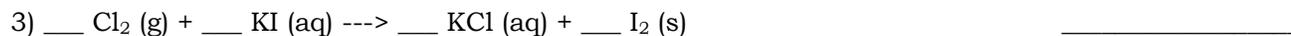


Equations Review

Name: _____

Part 1: Supply the correct coefficients to balance the following equations (assume all formulas are correct). Then classify each reaction as either synthesis, decomposition, single replacement, double replacement or combustion.

Type of Reaction



Part 2: Write the correct formulas for all reactants and products, then supply the necessary coefficients to balance the equations. Then classify each equation according to its reaction type.

1) Aqueous calcium iodide is mixed with aqueous mercury (II) nitrate to produce solid mercury (II) iodide and aqueous calcium nitrate.

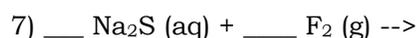
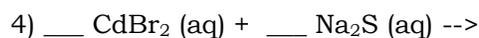
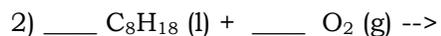
2) Aqueous hydrogen peroxide decomposes to form liquid water and oxygen gas.

3) Solid aluminum mixed with chlorine gas produces solid aluminum chloride.

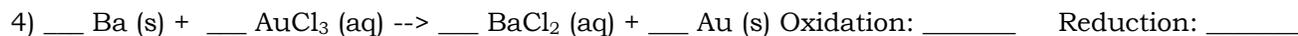
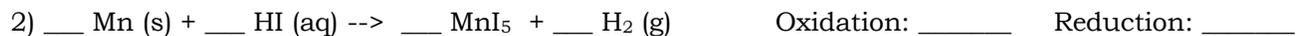
4) Manganese metal reacts with aqueous sulfuric acid (H_2SO_4) to produce aqueous manganese (II) sulfate and hydrogen gas.

5) Aqueous barium acetate reacts with aqueous sodium phosphide to produce solid barium phosphide and aqueous sodium acetate.

Part 3: Given the reactants, predict the products and balance the equation:



Part 4: In the following reactions, balance the reaction and determine what element undergoes oxidation and what element undergoes reduction:



Part 5: Using the activity series, determine whether the reaction will occur or not. For each reaction, predict the products and balance. If no reaction occurs, write NR.

