Molarity

1) What is the molarity of a solution which contains 2.65 moles of NaCl dissolved to a total volume of 600. mL?

2) What is the molarity of a solution prepared by dissolving 14.1 g of ammonia (NH_3) in water to a final volume of 0.100 liter?

3) How many moles of sodium fluoride would need to dissolve in 125 mL of water to make a 2.75 M solution?

4) How many grams of FeCl₃•6H₂O would be dissolved in 750. mL of 1.25 M solution?

5) How many grams of sodium hydroxide must be used to make 3.00 L of 0.450 M solution?

6) What volume (in mL) of 0.250 M solution could be made from 8.75 grams of potassium chloride?

7) How many grams of $CuSO_4 \bullet 5H_2O$ will be needed to make 75.0 mL of 0.500 M solution?

Molarity

Name: _

1) What is the molarity of a solution which contains 2.65 grams of calcium fluoride dissolved to a total volume of 600. mL?

2) What is the molarity of a solution prepared by dissolving 14.1 g of oxygen difluoride in water to a final volume of 0.100 liter?

3) How many moles of magnesium iodide would need to dissolve in 125 mL of water to make a 2.75 M solution?

4) How many grams of Ni(NO₃)₂•2H₂O would be dissolved in 750. mL of 2.15 M solution?

5) How many grams of aluminum hydroxide must be used to make 4.00 L of 0.350 M solution?

6) What volume (in mL) of 0.250 M solution could be made from 87.5 grams of potassium dichromate?

7) How many grams of $BaSO_4 \bullet 7H_2O$ will be needed to make 75.0 mL of 0.500 M solution?