

Molarity

Name: _____

1) What is the molarity of a solution which contains 3.11 moles of KCl dissolved to a total volume of 1200. mL?

2) What is the molarity of a solution prepared by dissolving 78.2 g of ammonium chloride in water to a final volume of 0.100 liter?

3) How many moles of barium iodide would need to dissolve in 125 mL of water to make a 2.75 M solution?

4) How many grams of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ would be dissolved in 750. mL of 1.25 M solution?

5) How many grams of rubidium hydroxide must be used to make 2.00 L of 0.150 M solution?

6) What volume (in mL) of 1.250 M solution could be made from 80.75 grams of nickel (II) chloride?

7) How many grams of $\text{CoCl}_2 \cdot 3\text{H}_2\text{O}$ will be needed to make 75.0 mL of 0.500 M solution?

Answers: 1) 2.59 M

4) 234 g

6) 498 mL