

Molecular Compounds

Name: _____

Part 1: For the following molecular compounds, if given the name, write the formula. If given the formula, write the name. Also, determine the polarity of the bonds.

Name	Formula	Bond Polarity	Name	Formula	Bond Polarity
	PCl ₃		Nitrogen monoxide		
	SF ₄		Phosphorous pentafluoride		
	SO ₃		Dinitrogen pentasulfide		
	H ₂ S		Carbon disulfide		
	S ₂ F ₁₀		Boron trichloride		
	Cl ₄		Tetraphosphorous trisulfide		
	N ₂ O		Chlorine dioxide		

Part 2: The following is a mixture of ionic and molecular compounds. If given the name, write the formula. If give the formula, write the name. It may be helpful to use the blank before the number to identify as ionic or molecular.

_____ 1) SiO₂

_____ 8) sodium thiocyanate

_____ 2) SbF₅

_____ 9) magnesium nitrate

_____ 3) CoWO₄

_____ 10) carbon monoxide

_____ 4) K₂Cr₂O₇

_____ 11) nitrogen trihydride

_____ 5) N₂O

_____ 12) iron (III) chloride

_____ 6) CCl₄

_____ 13) potassium cyanide

_____ 7) Na₃PO₄

_____ 14) silicon tetrafluoride