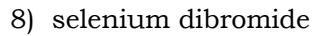
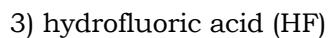
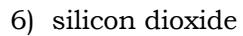
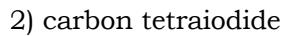
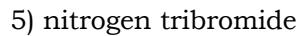
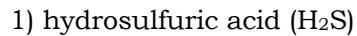


## Molecular Shape

Name: \_\_\_\_\_

Directions: For each of the following,

- a) Draw the electron dot (Lewis) diagram of the molecule.
  - b) Determine the bond type (polar or nonpolar).
  - c) Determine the shape of the molecule and
  - d) Determine whether the molecule is polar or nonpolar.
- \*Use your electronegativity chart to determine bond polarity.



Part 2: For each of the following:

- a) Draw the Lewis diagram for the following compounds.
- b) Determine whether the molecule is polar or nonpolar.
- c) Then, draw the molecule in it's proper shape using the structural drawing method.

1) water

5) carbon disulfide

2) selenium monoxide

6) silicon tetraiodide

3) oxygen difluoride

7) phosphorous trichloride

4) nitrogen trichloride

8) iodine monofluoride