

Stoichiometry 2

Name: _____

Directions: Write the balanced equation for each reaction and show all steps in your solution. Remember to include units in your work.

1) Tin (II) fluoride and hydrogen can be produced by the reaction of solid tin with hydrogen fluoride gas.

Equation:

a) How many grams of tin (II) fluoride can be produced by the reaction of 128 g of hydrogen fluoride?

b) How many molecules of hydrogen gas would be produced by the reaction?

2) Limestone (CaCO_3) can be decomposed into carbon dioxide and calcium oxide.

Equation:

a) How many liters of carbon dioxide can be produced, at STP, from 300 g of limestone?

b) If 78.3 L of carbon dioxide is produced from a sample of limestone, how many grams of calcium oxide was produced?

Answers: 1a) 502.4 g tin (II) fluoride

2b) 196 g calcium oxide