

Titration and Neutralization

Name: _____

1) Complete and balance the following neutralization reactions.



2) In a titration reaction, 43.0 mL of NaOH solution was required to neutralize 32.0 mL of 0.100 M HCl solution. What is the molarity of the NaOH solution?

3) The concentration of an acetic acid (HC₂H₃O₂) solution is determined by titrating 25.0 mL of the solution with KOH. If 42.5 mL of 0.075 M KOH is required to reach the equivalence point, what is the molarity of the acetic acid?

4) It requires 50.0 mL of 0.150 M NaOH solution to neutralize a 0.345 M HF solution. What volume of the hydrofluoric acid was neutralized?

5) In a titration, a 0.450 M solution of acetic acid is used to neutralize 34.0 mL of 0.125 M copper (I) hydroxide solution (CuOH). What is the volume of the acetic acid will need to be used?

7) Write formulas for the following acids and bases:

a) hydroselenic acid _____ e) strontium hydroxide _____

b) sulfous acid _____ f) periodic acid _____

c) chromium (III) hydroxide _____ g) cyanic acid _____

d) bromic acid _____ h) hydroarsenic acid _____

Answers: 1) 0.0744 M 3) 21.7 mL

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