

Writing Chemical Equations

Name: _____

For questions 1-3, write chemical equations for the following word reactions. For questions 4-6, write out the chemical equations in words.

1) Two aqueous sodium phosphate molecules plus three aqueous copper (II) sulfate molecules produces three aqueous sodium sulfate molecules and one solid copper (II) phosphate molecule.

2) Aqueous iron (III) oxide reacts with three aqueous calcium bromate to produce two solid iron (III) bromate and three aqueous calcium oxide.

3) Two solid magnesium can react with oxygen gas and make two solid magnesium oxide.

4) $\text{SnCO}_3 (\text{s}) + \text{H}_2\text{CrO}_4 (\text{aq}) \rightarrow \text{SnCrO}_4 (\text{s}) + \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{l})$

5) $\text{CuCl}_2 (\text{aq}) + \text{Mg} (\text{s}) \rightarrow \text{MgCl}_2 (\text{aq}) + \text{Cu} (\text{s})$

6) $2 \text{Fe} (\text{s}) + 3 \text{H}_2\text{O} (\text{l}) \rightarrow \text{Fe}_2\text{O}_3 (\text{s}) + 3 \text{H}_2 (\text{g})$

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1) Three solid magnesium reacts with two aqueous aluminum nitrate to make three aqueous magnesium nitrate and two solid aluminum.

2) Two solid sodium bicarbonate can be heated to form solid sodium oxide, liquid water and two gaseous carbon dioxide.

3) Three liquid bromine can be mixed with two solid chromium to make two solid chromium (III) bromide.

4) $\text{KBr (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{KNO}_3 \text{ (aq)} + \text{AgBr (s)}$

5) $(\text{NH}_4)_2\text{SO}_4 \text{ (aq)} + \text{MnCl}_2 \text{ (aq)} \rightarrow 2 \text{ NH}_4\text{Cl (aq)} + \text{MnSO}_4 \text{ (s)}$

6) $2 \text{ KClO}_3 \text{ (s)} \rightarrow 2 \text{ KCl (s)} + 3 \text{ O}_2 \text{ (g)}$