

Writing Chemical Equations

Name: _____

For questions 1-3, write chemical equations for the following word reactions. For questions 4-6, write out the chemical equations in words.

1) Two aqueous beryllium iodide molecules plus one aqueous tin (IV) nitrate molecule produces two aqueous beryllium nitrate molecules and one solid tin (IV) iodide molecule.

2) Two solid iron atoms plus three fluorine gas molecules yield two solid iron (III) fluoride.

3) Two aqueous ammonium nitrate molecules added to three aqueous lead (II) chlorate molecules react to give six aqueous ammonium chlorate molecules and one solid lead (II) nitrate molecule.

4) $\text{KBr (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{KNO}_3 \text{ (aq)} + \text{AgBr (s)}$

5) $\text{CuCl}_2 \text{ (aq)} + \text{Mg (s)} \rightarrow \text{MgCl}_2 \text{ (aq)} + \text{Cu (s)}$

6) $2 \text{ Fe (s)} + 3 \text{ H}_2\text{O (l)} \rightarrow \text{Fe}_2\text{O}_3 \text{ (s)} + 3 \text{ H}_2 \text{ (g)}$